

More Chapters 14 Study Questions

1. Calculate the pH of a 0.030 M NH_4Cl solution.
2. Calculate the pH of a 0.14 M solution of NaF.
3. Find the pH of the following solution: 6.00 mL of 3.00 M HNO_3 diluted to a volume of 18.0 liters with water.
4. Write a balanced *net ionic* equation for the reaction between NH_4^+ and NaOH. What are the conjugate acid/base pairs in this reaction?
5. Which of the following has the highest pH?
a) 0.10 M NaCl b) 0.10 M HCl c) 0.10 M NaOH d) 0.10 M NaNO_2