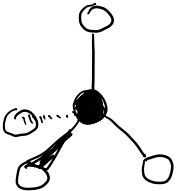
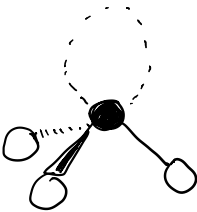
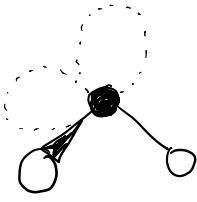
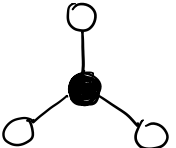
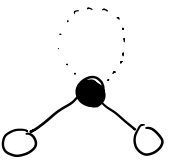



MOLECULAR GEOMETRY

#bonding groups	# bonds	# lone pairs	Structure	Electron Geometry	Molecular Geometry	Example	Lewis Structure
4	4	0		Tetrahedral	Tetrahedral	CH ₄	$\begin{array}{c} \text{H} \\ \\ \text{H}-\text{C}-\text{H} \\ \\ \text{H} \end{array}$
4	3	1		Tetrahedral	Trigonal pyramid	NH ₃	$\begin{array}{c} \text{H}-\ddot{\text{N}}-\text{H} \\ \\ \text{H} \end{array}$
4	2	2		Tetrahedral	Bent	H ₂ O	$\text{H}-\ddot{\text{O}}-\text{H}$

#bonding groups	# bonds	# lone pairs	Structure	Electron Geometry	Molecular Geometry	Example	Lewis Structure
3	3	0		Trigonal planar	Trigonal planar	CH ₂ O	$\begin{array}{c} \text{H} \\ \diagdown \\ \text{C}=\ddot{\text{O}} \\ \diagup \\ \text{H} \end{array}$
3	2	1		Trigonal planar	Bent	SO ₂	$\begin{array}{c} \text{O} \\ \diagdown \\ \text{S}=\ddot{\text{O}} \\ \diagup \\ \text{O} \end{array}$

#bonding groups	# bonds	# lone pairs	Structure	Electron Geometry	Molecular Geometry	Example	Lewis Structure
2	2	0		Linear	Linear	CO ₂	$\ddot{\text{O}}=\text{C}=\ddot{\text{O}}$